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Der Krokodill und das Meerpferd.

Einst lag das Unthier an dem Strande
Des Stroms gestreckt, und dörrte sich
Den feuchten Ranzgen säuberlich
In der Aegypter heißem Sande.
Ein armes Kind, das noch nicht viel
Von diesem Ungeheuer wußte,
Und sich dem Flusse nähern mußte,
Kam aus Versehen zum Krokodill.

Ach, dieses hab' ich längst gewünscht,
 Daß dich der Himmel annoch liebte,
 Hüb dieser Meerapostel an;
 Bedauere du nur dein Verbrechen,
 Und weine ferner Tag und Nacht,
 Daß du dieß Kinklein umgebracht,
 So wird die Vorsecht es nicht rächen.
 Da wär' ich so ein Thor, wie du,
 Schrie ihm der Neubekehrte zu,
 Erspare künftig deine Lehren;
 Der Zunge macht mich noch nicht satt,
 Weil er kein Fleisch am Kopfe hat,
 Das ist die Ursach' meiner Zähren.

Ihr frommen Seelen, traut des Heuchlers Thränen nicht,
Denn was er mit dem Munde spricht,
Das läugnet er in seinem Herzen,
Sein Auge weint, und die Gedanken scherzen.

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VI.

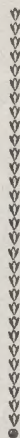
Der Fuchs und der Adler.



Es lebt' aus Reineckens Geschlechte
Ein jung und eitler Abkömmling,
Der oft mit mehrerm Glück als Rechte
Der schnellen Hunde Spur entging.

Da lag er nun vor seinem Loche,
Und lachte bey sich der Gefahr,
Der er noch in vergangner Woche
Durch einen Sprung entronnen war.

Sagt, rief er, Hüfe, Wiesen, Ställe,
Ihr Zeugen meiner Tapferkeit,
Wer stiehlt, wie ich? Wer sieht so helle?
Wer läuft so schnell? Wer riecht so weit?



Vertieft in solchen Wunderdingen
Bemerkt' er eines Adlers Flug,
Wie ihn mit ausgestreckten Schwingen
Das stille Meer der Lüfte trug.

O könnt' ich fliegen, wie die Vögel!
Den Reid, erseufzt' er, macht' ich stumm,
Euch aber kahl, ihr Bauernflegel;
Mit Lust gäb' ich ein Ohr darum.

Ist legt ein Schuß den Adler nieder,
Der Fuchs nimmt es mit Schrecken wahr,
Zu fliegen wünscht er nimmer wieder:



Je höher Stand, je mehr Gefahr.

[illegible]

1. The first part of the paper is devoted to a discussion of the
 2. various methods which have been employed for the determination of
 3. the rate of reaction between a gas and a solid. It is shown that
 4. the most reliable method is that of measuring the change in weight
 5. of the solid during the reaction. This method is applicable to all
 6. cases in which the reaction is accompanied by a change in weight.
 7. The second part of the paper is devoted to a discussion of the
 8. various factors which influence the rate of reaction between a gas
 9. and a solid. It is shown that the rate of reaction is influenced
 10. by the nature of the solid, the nature of the gas, the temperature
 11. of the reaction, and the surface area of the solid. The third
 12. part of the paper is devoted to a discussion of the various
 13. theories which have been proposed to explain the rate of reaction
 14. between a gas and a solid. It is shown that the most satisfactory
 15. theory is that of the collision theory. This theory states that the
 16. rate of reaction is proportional to the number of collisions
 17. between the gas molecules and the solid surface. The fourth
 18. part of the paper is devoted to a discussion of the various
 19. applications of the collision theory to the rate of reaction between
 20. a gas and a solid. It is shown that the collision theory can be
 21. applied to all cases in which the reaction is accompanied by a
 22. change in weight. The fifth part of the paper is devoted to a
 23. discussion of the various factors which influence the rate of
 24. reaction between a gas and a solid. It is shown that the rate of
 25. reaction is influenced by the nature of the solid, the nature of the
 26. gas, the temperature of the reaction, and the surface area of the
 27. solid. The sixth part of the paper is devoted to a discussion of
 28. the various theories which have been proposed to explain the rate
 29. of reaction between a gas and a solid. It is shown that the most
 30. satisfactory theory is that of the collision theory. This theory
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 32. collisions between the gas molecules and the solid surface. The
 33. seventh part of the paper is devoted to a discussion of the various
 34. applications of the collision theory to the rate of reaction between
 35. a gas and a solid. It is shown that the collision theory can be
 36. applied to all cases in which the reaction is accompanied by a
 37. change in weight. The eighth part of the paper is devoted to a
 38. discussion of the various factors which influence the rate of
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 40. reaction is influenced by the nature of the solid, the nature of the
 41. gas, the temperature of the reaction, and the surface area of the
 42. solid. The ninth part of the paper is devoted to a discussion of
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 48. tenth part of the paper is devoted to a discussion of the various
 49. applications of the collision theory to the rate of reaction between
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 51. applied to all cases in which the reaction is accompanied by a
 52. change in weight. The eleventh part of the paper is devoted to a
 53. discussion of the various factors which influence the rate of
 54. reaction between a gas and a solid. It is shown that the rate of
 55. reaction is influenced by the nature of the solid, the nature of the
 56. gas, the temperature of the reaction, and the surface area of the
 57. solid. The twelfth part of the paper is devoted to a discussion of
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 60. satisfactory theory is that of the collision theory. This theory
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 62. collisions between the gas molecules and the solid surface. The
 63. thirteenth part of the paper is devoted to a discussion of the
 64. various applications of the collision theory to the rate of reaction
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 67. by a change in weight. The fourteenth part of the paper is
 68. devoted to a discussion of the various factors which influence the
 69. rate of reaction between a gas and a solid. It is shown that the
 70. rate of reaction is influenced by the nature of the solid, the
 71. nature of the gas, the temperature of the reaction, and the
 72. surface area of the solid. The fifteenth part of the paper is
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 78. molecules and the solid surface. The sixteenth part of the paper
 79. is devoted to a discussion of the various applications of the
 80. collision theory to the rate of reaction between a gas and a solid.
 81. It is shown that the collision theory can be applied to all cases
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 83. seventeenth part of the paper is devoted to a discussion of the
 84. various factors which influence the rate of reaction between a gas
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 88. eighteenth part of the paper is devoted to a discussion of the
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 93. collisions between the gas molecules and the solid surface. The
 94. nineteenth part of the paper is devoted to a discussion of the
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 98. by a change in weight. The twentieth part of the paper is
 99. devoted to a discussion of the various factors which influence the
 100. rate of reaction between a gas and a solid. It is shown that the
 101. rate of reaction is influenced by the nature of the solid, the
 102. nature of the gas, the temperature of the reaction, and the
 103. surface area of the solid. The twenty-first part of the paper is
 104. devoted to a discussion of the various theories which have been
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 113. cases in which the reaction is accompanied by a change in weight.
 114. The twenty-third part of the paper is devoted to a discussion of
 115. the various factors which influence the rate of reaction between a
 116. gas and a solid. It is shown that the rate of reaction is
 117. influenced by the nature of the solid, the nature of the gas, the
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 125. surface. The twenty-fifth part of the paper is devoted to a
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 134. reaction, and the surface area of the solid. The twenty-seventh
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 142. various applications of the collision theory to the rate of reaction
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 144. can be applied to all cases in which the reaction is accompanied
 145. by a change in weight. The twenty-ninth part of the paper is
 146. devoted to a discussion of the various factors which influence the
 147. rate of reaction between a gas and a solid. It is shown that the
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 149. nature of the gas, the temperature of the reaction, and the
 150. surface area of the solid. The thirtieth part of the paper is
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 161. The thirty-second part of the paper is devoted to a discussion
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 165. temperature of the reaction, and the surface area of the solid.
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 190. a solid. It is shown that the collision theory can be applied to
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 192. weight. The thirty-eighth part of the paper is devoted to a
 193. discussion of the various factors which influence the rate of
 194. reaction between a gas and a solid. It is shown that the rate of





Der Fuchs und das Eichhorn.

Das muß ich wohl mit Dank erkennen,
 Verseht das Eichhorn, daß du mich
 So heftig liebst, ich bitte dich,
 Kannst du mir deinen Namen nennen?
 Zu dienen, Eichhorn heißet er,
 Dein Vater, tröst ihn Jupiter,
 Und meiner, waren rechte Brüder,
 Völbürge Brüder, und wir sind
 Im andern Grad gesippt, mein Kind!
 O steige doch geschwind hernieder.

So sind wir zwey so nahe Vettern,
Antwortete das Eichhorn drauf,
So werd' ich, nimm's nicht übel auf,
Annoch ein wenig höher klettern.
Denn meine Mutter lehrte mich,
Daß unter nahen Vettern sich
Die Eintracht allzeit stärker nähre,
Je weiter hier auf dieser Welt,
Wo Mein und Dein uns fallen stellt,
Der eine von dem andern wäre.

Der gute Fuchs ging seine Straße,
Und dachte, daß der Unterricht
Von seiner alten Ruhme nicht
Auf all' und jede Fälle passe,
Nur dieses fiel, mit allem dem,
Dem alten Heuchler unbecquem,
Daß sein Gewissen ihn belehrte,
Daß unter die, bey denen man
Die Lehre wirklich brauchen kann,
Er und sein Vetter auch gehörte.

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Die z w e y H ä h n e.

Sie rüßten auf einander los,
Den Zweykampf muthig anzuhoben,
Wer gibt, wer kriegt den ersten Stoß?
Wer unter beyden läßt das Leben?

Kopf gegen Kopf, Hahn gegen Hahn,
Sieht man im kurzen Lager stehen,
Sie sehn sich ein halb Stündchen an,
Da sie still aus einander gehen.

Sie blieben ehrlich nach wie vor,
Die kurze Thorheit ist die beste,
Wer Zweykampf sucht, der ist ein Thor,
Und wer sich schlägt, der ist der größte.

Es ist ein Kunst, die man nicht
 ohne viel Mühe und Arbeit
 lernen kann. Man muß
 sich viel anstrengen, um
 sie zu erlangen.

Die Kunst des Schreibens
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